

Chapter 13 States Of Matter



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The electron density around each nucleus is, for a moment, greater in one region of each cloud. Each molecule forms a temporary dipole. 394 Chapter 13 States of Matter. When temporary dipoles are close together, a weak dispersion force exists between oppositely charged regions of the dipoles, as shown in Figure 13-8.

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Chapter 13 States of Matter - Chapter 13 "States of... Section 13.2 The Nature of Liquids OBJECTIVES: Describe the equilibrium between a liquid and its vapor. Section 13.2 The Nature of Liquids OBJECTIVES: Identify the conditions at which boiling occurs. Section 13.2 The Nature of Liquids Liquid particles are also in motion.

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Chapter 13 States of Matter 137. SECTION 13.1 THE NATURE OF GASES (pages 385–389) This section introduces the kinetic theory and describes how it applies to gases. It defines gas pressure and explains how temperature is related to the kinetic energy of the particles of a substance. Kinetic Theory and a Model for Gases (pages 385–386)

Name Date Class STATES OF MATTER 13

13.1 The Fluid States 300 States of Matter FIGURE 13–1 The ice cube, a solid, has a definite shape. But water, a fluid, takes the shape of its container.

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Chapter 13 States of Matter notes. So $F/A = F'/A'$. A small force applied over a small area can give a bigger force over a bigger area like in a hydraulic system. The pressure you feel under water is due to the weight of the water over you. For a uniform fluid this depends on the density and the depth. $P = \rho hg$ where ρ is the density of the fluid.

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12 Best Images of States Of Matter Worksheet Answer Key ...

Chapter 13- The States of Matter Gases- indefinite volume and shape, low density. Liquids- definite volume, indefinite shape, and high density. Solids- definite volume and shape, high density Solids and liquids have high densities because their molecules are close together.

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13 STUDY GUIDE FOR CONTENT MASTERY CHAPTER States of Matter Section 13.1 Gases In your textbook, read about the kinetic-molecular theory. Complete each statement. 1. The kinetic molecular theory describes the behavior of gases in terms of particles in 2. The kinetic-molecular theory makes the following assumptions. a.

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graph representing the relationships among the solid, liquid, and vapor states of a substance in a sealed container standard atmosphere pressure required to support 760 mm of mercury at 25°C

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Chapter 13 - States of Matter - 13.3 The Nature of Solids - 13.3 Lesson Check - Page 434: 24 Answer Solids have a definite shape and a definite volume because of the arrangement of the particles.

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Chemistry Chapter 13: States Of Matter; Erin C. • 67 cards. Kinetic - Molecular Theory. All matter consists of tiny particles that are in constant motion. Kinetic - Molecular Theory. Used to explain the properties of matter in terms of the energy and movement of the particles and the forces acting between them. ...

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